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*Net Ionic*

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**Precipitation**

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*Solubility Rules*

*and How to Use a*

*Solubility Table*

*What Happens*

*when Stuff*

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*Reactions*

*Aqueous vs.*

*Solid: Chem*

*Class Solubility*

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Partial

Pressures \u0026amp;

Vapor Pressure:

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Massive

Creatures: Crash

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4.1-4.4

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Ionic Equations,  
and Chemical

~~Changes~~ Chemical

~~Reactions in~~

~~Aqueous~~

~~Solutions — Part~~

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Reactions in

Aqueous

Solutions - Part

II *Water* \u0026amp;

*Solutions - for*

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*Dirty Laundry:*

*Crash Course*

*Chemistry #7*

*Identifying*

*Strong*

*Electrolytes,*

*Weak*

*Electrolytes,*

*and*

*Nonelectrolytes*

*- Chemistry*

*Examples*

*Chemical*

*Reactions in*

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Reactions and

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Stoichiometry -

2 - Ionic

Compounds in

Water • When an

ionic compound

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dissolves in water, the ions are said to dissociate. •

This means that in solution, the solid no longer exists as a well-ordered arrangement of ions in contact ap chemistry ...

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proportion, and  
quantity;  
structure and  
properties of  
substances;  
transformations;  
and energy.

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Compounds in  
Water • When an  
ionic compound  
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water ...

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4.1 General

Properties of

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Solutions • A

solution is a

homogeneous

mixture of two

or more

substances. • A

solution is made

when one

substance (the

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(solute) is dissolved in another (the solvent). • The solute is the substance that is present in ...

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Calculate the

concentration of

urea, in mol/L,

in the saturated

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solution at  
20. °C. (d) The  
student also  
determines that  
the  
concentration of  
urea in a  
saturated  
solution at 25°C  
is 19.8 . M.

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2. Placing aluminum foil in any solution containing aqueous copper(II) ions will result in a reaction. The reaction is slow

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to begin with,  
then proceeds  
rapidly. a)

Referring to a  
solubility  
table, name at  
least four ionic  
compounds that  
could be  
dissolved in  
water to make a  
solution  
containing  
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copper(II) ions.

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