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Modern Chemistry • CHAPTER 9 HOMEWORK 9-2 (pp. 280-282) VOCABULARY ... Circle the letter of the best answer. 1. Given the balanced equation  $2C_4H_{10} + 13O_2 \rightarrow 8CO_2 + 10H_2O$ , how many moles of  $CO_2$  are produced when 14.9 g of  $O_2$  ... Microsoft Word - Homework 09-01 Author: Susan M. Free

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the following equation:  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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CHAPTER 6 REVIEW Chemical Bonding SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. a The notation for sodium chloride,  $NaCl$ , stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) ...

## 6-Chemical Bonding

Modern Chemistry Chapter 9 Stoichiometry. Sections 1-3 Introduction to Stoichiometry Ideal. Stoichiometric Calculations Limiting Reactant and. Percent Yield. 2. Definitions. Composition stoichiometry deals with the mass. relationships of elements in compounds. Reaction stoichiometry involves the mass.

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